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ion F a  $6.7 \times 10^{-4}$  b  $1.5 \times 10^{-3}$  c  $1.5 \times 10^{-11}$  d'

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Determine the  $\text{OH}^-$  and the  $\text{H}^+$  of a 0.20 M solution of formic acid  
The  $K_a = 1.8 \times 10^{-4}$   $\text{HCOOH} \rightleftharpoons \text{H}^+ + \text{HCOO}^-$

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